

Please check the examination details below before entering your candidate information

Candidate surname		Other names	
Pearson Edexcel International Advanced Level	Centre Number	Candidate Number	
Thursday 21 January 2021			
Afternoon (Time: 1 hour 20 minutes)		Paper Reference WCH16/01	
Chemistry International Advanced Level Unit 6: Practical Skills in Chemistry II			
You must have: Scientific calculator			Total Marks

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*
- Show all your working in calculations and include units where appropriate.

Information

- The total mark for this paper is 50.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- You will be assessed on your ability to organise and present information, ideas, descriptions and arguments clearly and logically, including your use of grammar, punctuation and spelling.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

P67820A

©2021 Pearson Education Ltd.

1/1/1/1/1




Pearson

Answer ALL the questions. Write your answers in the spaces provided.

- 1** A student carries out some tests on four aqueous solutions **A**, **B**, **C** and **D**.
One of the solutions is aqueous barium chloride, $\text{BaCl}_2(\text{aq})$.

- (a) The student is asked to add **A** to samples of **B**, **C** and **D** in separate test tubes, a **small** amount at a time, until there is no further change.

The container of solution **A** has a hazard label.



- (i) Identify the hazard indicated by this label.

(1)

- (ii) Describe how you would add small amounts of **A** until there is no further change. Name the apparatus you would use.

(2)



- (b) (i) **B** is a blue solution. When **A** is added to **B**, the mixture first turns green and then gradually turns yellow.

Give the **formula** of the cation in **B**.

(1)

- (ii) When **A** is added to **C**, vigorous effervescence occurs and the gas produced turns limewater cloudy.

Identify, by name or formula, the gas produced.

(1)

- (iii) Suggest the identity, by name or formula, of the anion in **C**.

(1)

- (iv) Identify **A** by name or formula. Justify your answer.

(2)

- (v) When **A** is added to **D** no change is seen.

A small amount of this mixture is added to **B** and a white precipitate forms.

Suggest what can be deduced about solutions **B** and **D**.

(2)

Solution **B**

Solution **D**



- (vi) A concentrated solution of ammonia is added to **B**.
Initially a pale blue precipitate forms. When more ammonia is added,
the precipitate dissolves forming a dark blue solution **F**.

Identify, by name or formula, the pale blue precipitate and the species
responsible for the dark blue colour in **F**.

(2)

- (vii) A solution of the sodium salt of EDTA, Na_4EDTA , is added to a sample of
solution **F**. The solution turns pale blue.

Write an equation for the reaction.
State symbols are not required.

(2)

(Total for Question 1 = 14 marks)



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

BLANK PAGE



P 6 7 8 2 0 A 0 5 1 6

- 2 Students were told to determine the concentration of a solution of potassium chlorate(V), KClO_3 . Two methods were used: precipitation and titration.

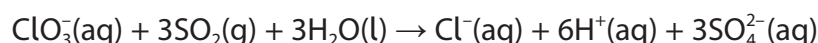
Method 1 – Precipitation

Step 1 Bubble excess sulfur dioxide, SO_2 , into 100 cm^3 of the potassium chlorate(V) solution.

Step 2 Boil the resulting mixture to remove excess SO_2 and then add silver nitrate solution until no more silver chloride precipitate forms.

Step 3 Filter, dry and weigh the precipitate.

The equation for the reaction in Step 1 is shown.



- (a) Identify the main hazard in Step 1, giving a safety precaution that will reduce the risk.

Assume that safety spectacles and a laboratory coat were used.

(2)

- (b) The reaction in Step 2 produced 0.430 g of a white precipitate of silver chloride, AgCl .

Calculate the concentration of KClO_3 in the solution, in mol dm^{-3} , found using Method 1.

You **must** show your working.

(2)



- (c) A student who used Method 1 obtained a value that was significantly larger than the actual concentration of the solution.

Explain **one** possible source of experimental error which might lead to this result.

(2)

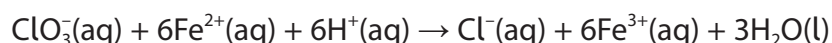
Method 2 – Titration

Step 1 Mix a sample of potassium chlorate(V) solution with an acidified solution containing iron(II) sulfate, FeSO_4

Step 2 Remove the chloride ions produced in Step 1.

Step 3 Determine the concentration of excess iron(II) ions by titrating the whole of the solution with a standard solution of potassium manganate(VII).

The equation for the reaction in Step 1 is shown.



- (d) Give the colour change observed in Step 1.

(1)



(e) Describe how to carry out the titration in Step 3. You should identify suitable apparatus and any additional chemicals required.

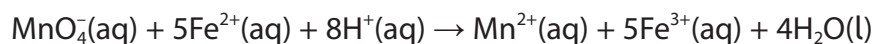
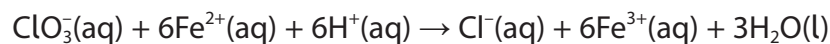
(5)



- (f) In Method 2, 50.0 cm³ of potassium chlorate(V) was mixed with 150 cm³ of 0.0750 mol dm⁻³ of iron(II) sulfate. The iron (II) sulfate was in excess.

The whole of this solution required 9.25 cm³ of 0.050 mol dm⁻³ of potassium manganate(VII) to completely react.

The equations for the reactions are



Calculate the concentration, in mol dm⁻³, of the potassium chlorate(V) solution. You **must** show your working.

(6)



- (g) Explain the change, if any, to the value calculated in (f) if the chloride ions were not removed before the reaction in Step 3 of Method 2.

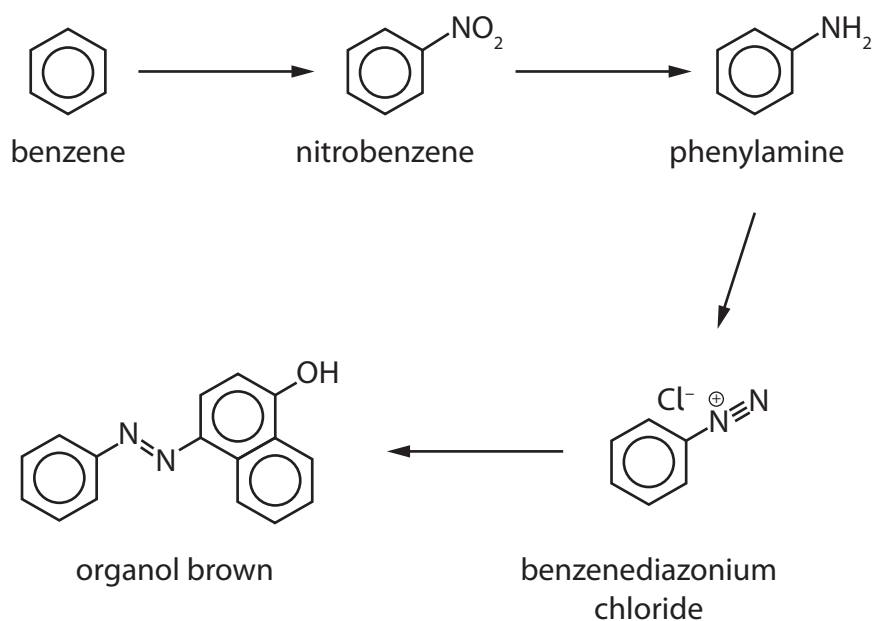
(2)

(Total for Question 2 = 20 marks)



- 3 Azo dyes, such as Organol Brown, can be made from benzene, C_6H_6 , using the reaction scheme shown.

Due to the toxicity of benzene, the first step is never carried out in a school laboratory.



- (a) In the preparation of nitrobenzene, benzene is added slowly to a mixture of concentrated nitric and sulfuric acids.

The mixture is warmed at $55^\circ C$ under reflux for 45 minutes. The reaction mixture is stirred continuously.

- (i) State why a reflux condenser is needed when the mixture is warmed.

(1)



(ii) Draw a diagram of the apparatus used to warm under reflux in this experiment.

(3)

(iii) Suggest why the reaction mixture is stirred continuously.

(2)

.....

.....

.....

.....

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



- (b) The excess acid is removed from the reaction mixture. The layer containing nitrobenzene is separated and dried before being purified by distillation.

Identify a suitable drying agent.

(1)

- (c) Nitrobenzene is then reduced to phenylamine, $C_6H_5NH_2$.

Phenylamine reacts with nitrous acid at a temperature between $0^\circ C$ and $10^\circ C$ to form a diazonium compound.

- (i) Nitrous acid is formed in the reaction mixture using sodium nitrite and hydrochloric acid.

State why nitrous acid is generated in the reaction mixture instead of being obtained from a chemical supplier.

(1)

- (ii) Explain why the temperature of the reaction between phenylamine and nitrous acid must be neither lower than $0^\circ C$ nor higher than $10^\circ C$.

(2)



- (d) Reaction of the diazonium compound with an alkaline solution of naphthalene-1-ol produces the solid azo dye, Organol Brown. The solid is purified by recrystallisation.

Procedure

Step 1 The impure Organol Brown is dissolved in a minimum volume of hot solvent.

Step 2 The solution is filtered hot through a preheated funnel.

Step 3 The solution is cooled and filtered using a Buchner funnel.

Step 4 The solid is rinsed with a small amount of ice-cold solvent.

Step 5 The solid is dried in a desiccator.

- (i) State why a **minimum** volume of hot solvent is used in Step 1.

(1)

- (ii) Explain why a preheated funnel is used in Step 2.

(1)

- (iii) Give a reason for each of the two filtrations in Steps 2 and 3.

(2)

- (iv) Give a possible reason why it is preferable to dry the solid in a desiccator rather than in an oven in Step 5.

(1)



- (e) The melting temperature of the recrystallised Organol Brown is measured to check its purity.

State what you would observe if the sample was pure.

(1)

(Total for Question 3 = 16 marks)

TOTAL FOR PAPER = 50 MARKS



16

Elements with atomic numbers 112-116 have been reported but not fully authenticated

* Actinide series

140	Ce	cerium	58	141	Pr	praseodymium	59	144	Nd	neodymium	60	147	Pm	promethium	61	150	Sm	samarium	62	152	Eu	europium	63	157	Gd	gadolinium	64	159	Tb	terbium	65	163	Dy	dysprosium	66	165	Ho	holmium	67	167	Er	erbium	68	169	Tm	thulium	69	173	Yb	ytterbium	70	175	Lu	lutetium	71
232	Th	thorium	90	231	Pa	protactinium	91	238	U	uranium	92	237	Np	neptunium	93	242	Pu	plutonium	94	243	Am	americium	95	247	Cm	curium	96	245	Bk	berkelium	97	251	Cf	californium	98	254	Es	einsteinium	99	253	Fm	fermium	100	256	Md	mercurium	101	254	No	nobelium	102	257	Lr	lawrencium	103

DO NOT WRITE IN THIS AREA

